

$$\textcircled{2d} \quad 1:1,000 \text{ epinephrine} = \frac{1 \text{ g epinephrine}}{1,000 \text{ mL}}$$

$$\Rightarrow \frac{1,000 \text{ mg}}{1,000 \text{ mL}} = \boxed{1 \text{ mg/mL}}$$

$$\textcircled{3a} \quad .15 \text{ M NaCl} = \frac{.15 \text{ mole NaCl}}{\text{L H}_2\text{O}}$$

$$1 \text{ mole of NaCl weighs} = 58.5 \text{ g}$$

$$.15 \text{ mole of NaCl weighs} = .15 \times 58.5 \text{ g}$$

$$= 8.775 \text{ g}$$

$$\approx 9 \text{ g}$$

$$= \boxed{\frac{9 \text{ g NaCl}}{1 \text{ L H}_2\text{O}}}$$

$$\textcircled{3b} \quad \frac{9 \text{ g NaCl}}{1,000 \text{ mL H}_2\text{O}} \rightarrow \frac{x}{10 \text{ mL}}$$

$$1,000 x = 90$$

$$x = \frac{90}{1,000} = .09 \text{ g}$$

$$\boxed{\frac{.09 \text{ g NaCl}}{10 \text{ mL H}_2\text{O}}}$$