

$$\textcircled{3c} \quad .15 \text{ M NaCl} = \frac{9 \text{ g NaCl}}{1,000 \text{ mL H}_2\text{O}}$$

(see 3a)

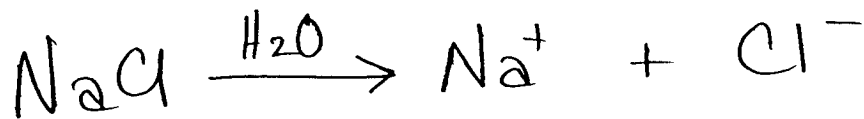
$$\frac{9 \text{ g}}{1,000 \text{ mL}} \xrightarrow{\text{X}} \frac{\text{X}}{100 \text{ mL}}$$

$$1,000 \text{ X} = 900$$

$$\text{X} = \frac{900}{1,000} = .9$$

$$= \frac{.9 \text{ g}}{100 \text{ mL}} = \boxed{.9\% \text{ NaCl}}$$

$$\textcircled{3d} \quad .15 \text{ M NaCl} = \frac{.15 \text{ mole NaCl}}{1 \text{ L H}_2\text{O}}$$



.15

x 2

= .3 osmole

$$= \boxed{\frac{300 \text{ milliosmoles}}{\text{L}}}$$